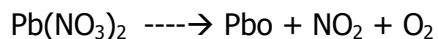


I. Very short questions:

1. Balance the following equations



2. Name the acid present in vinegar.

3. Give one example of a reaction which is a double displacement reaction as well as a precipitation reaction.

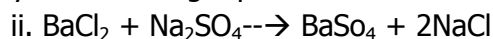
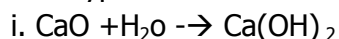
4. Define oxidation and reduction.

5. Why does dry HCl gas not change the colour of the dry litmus paper?

II. Short answer questions.

1. In the reaction $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$, name the substance which is oxidised and which is reduced.

2. Name the type of chemical reaction represented by the following equation.



3. List two differences between an acid and a base based on their chemical properties.

4. While diluting an acid, why is it recommended that acid should be added to water and not water to that?

5. Translate the following statements into chemical equations.

i. Magnesium burns in the presence of nitrogen to form magnesium nitride.

ii. Lead nitrate reacts with sulphuric acid to form a precipitate of lead sulphate and nitric acid.

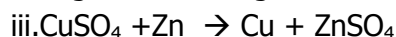
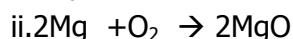
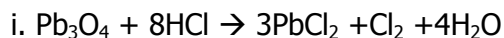
III. Long answer question.

1.a. Which gas is usually liberated when an acid reacts with a metal? Illustrate with an example. How will you test for the presence of this gas?

b. Why do we apply paint on iron articles?

2.a. Why should curd and sour substance not be kept in brass and copper vessels?

b. Identify the oxidising agent in the following reaction.



3. Write the balanced chemical equation for the following and identify the type of reaction in each case.

1. Potassium bromide + Barium iodide \rightarrow Potassium iodide + Barium bromide2. Magnesium + Hydrochloric acid \rightarrow Magnesium Chloride + Hydrogen3. Zinc carbonate \rightarrow zinc oxide + carbon dioxide